



# Standard Test Methods for Silanes Used in Rubber Formulations (bis-(triethoxysilylpropyl)sulfanes): Sulfur Content<sup>1</sup>

This standard is issued under the fixed designation D6741; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon ( $\epsilon$ ) indicates an editorial change since the last revision or reapproval.

## 1. Scope

1.1 These test methods cover the determination of the sulfur content of silanes, or of admixtures of silane and carbon black, of the type bis-(triethoxysilylpropyl)sulfane. The following test methods are included:

- Test Method A: Combustion in an Oxygen Flask and Subsequent Titration of Sulfate
- Test Method B: High-Temperature Combustion with Infrared Absorption Detection

1.2 The values stated in SI units are to be regarded as the standard. The values given in parentheses are for information only.

1.3 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.*

## 2. Referenced Documents

2.1 *ASTM Standards:*<sup>2</sup>

- [D1193 Specification for Reagent Water](#)
- [E177 Practice for Use of the Terms Precision and Bias in ASTM Test Methods](#)
- [E691 Practice for Conducting an Interlaboratory Study to Determine the Precision of a Test Method](#)

## 3. Significance and Use

3.1 The sulfur content can be used to characterize a silane or an admixture of silane and carbon black. It depends mainly on the average length of the sulfur chain.

<sup>1</sup> These test methods are under the jurisdiction of ASTM Committee D11 on Rubber and are the direct responsibility of Subcommittee D11.20 on Compounding Materials and Procedures.

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<sup>2</sup> For referenced ASTM standards, visit the ASTM website, [www.astm.org](http://www.astm.org), or contact ASTM Customer Service at [service@astm.org](mailto:service@astm.org). For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

## Test Method A: Combustion in an Oxygen Flask and Subsequent Titration of Sulfate

### 4. Summary of Test Method

4.1 The weighed sample is burned in an oxygen atmosphere. The resulting sulfur dioxide and trioxide is absorbed in hydrogen peroxide solution, and the sulfate ions thus generated are titrated with barium perchlorate solution employing Thorin indicator. Taking into account the known concentration of the perchlorate solution, the volume consumed, and the sample mass, the sulfur content of the sample can be calculated.

### 5. Apparatus

- 5.1 *Analytical Balance*, accuracy  $\pm 0.1$  mg.
- 5.2 *Combustion Apparatus*, according to Schoeniger, with 750 cm<sup>3</sup> flask.
- 5.3 *Ultrasonic Bath*.
- 5.4 *Automatic Titrator*.
- 5.5 *Photoelectrical Detector*, set at 550 nm.
- 5.6 *Volumetric Cylinder*, 100 cm<sup>3</sup>.
- 5.7 *Volumetric Flask*, 100 cm<sup>3</sup>.
- 5.8 *Volumetric Pipet*, 10 cm<sup>3</sup>.
- 5.9 *Syringe*, 0.5 cm<sup>3</sup>.
- 5.10 *Titration Beakers*.
- 5.11 *Magnetic Stirrer*, with stirring rod.
- 5.12 *Filter Paper*, for Schoeniger combustion (35 × 35 × 70 mm).
- 5.13 *Glass Tube*, attached to a rubber hose.

### 6. Reagents

6.1 *Purity of Reagents*—Reagent grade shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society where

such specifications are available.<sup>3</sup> Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

6.2 *Purity of Water*—Unless otherwise indicated, references to water shall be understood to mean reagent water conforming to Specification **D1193**.

6.3 *2-propanol* > 99 %.

6.4 *Hydrogen Peroxide Solution*, (5 %) as absorbing liquid.

6.5 *Barium Perchlorate Solution*, 0.005 mol/dm<sup>3</sup> in 2-propanol/water.

6.6 *Thorin Solution*, 0.1 % in water (Thorin octahydrate).

6.7 *Sulfuric Acid*, as a standard solution (1.000 ± 0.002 g SO<sub>4</sub><sup>2-</sup>).

6.8 *Deionized Water*.

6.9 *Oxygen*.

6.10 *Sulfanilic Acid*, (certified standard for elemental analysis).

## 7. Preparation of Apparatus

NOTE 1—The flask shall not contain any residues from solvents. Purge with compressed air if necessary.

NOTE 2—Check flask for damages. Do not use flasks having scratches or cracks.

NOTE 3—Ignite with safety shield lowered only.

NOTE 4—After combustion is finished, leave flask in the apparatus for 5 min cool-down.

7.1 Add 20 cm<sup>3</sup> of 5 % H<sub>2</sub>O<sub>2</sub> solution into the combustion flask by means of a volumetric cylinder.

7.2 Add a magnetic stirring rod.

7.3 Purge the flask for 1 min with oxygen gas. Do this by directing a gentle stream from the glass tube towards the bottom of the flask.

## 8. Procedure

8.1 *Preparation of Liquid Samples (Silanes)*:

8.1.1 Place a non-densified cotton wool swab between the platinum grids attached to the glass stopper. The spare place between the grids should be filled to  $\frac{2}{3}$ .

8.1.2 Attach the ignition strip in such way that it is in contact with the heating coil.

8.1.3 Transfer the stopper onto the balance, add approximately 60 mg of silane and record the mass to the nearest 0.1 mg.

8.2 *Preparation of Solid Samples (Admixtures of Carbon Black and Silane)*:

8.2.1 Fold the special filter paper as to form a tunnel and transfer it onto the balance.

8.2.2 Add approximately 120 mg of the admixture and record the mass to the nearest 0.1 mg.

8.2.3 Fold the paper, roll it together and insert it between the platinum grids. Make sure the ignition strip is in contact with the heating coil.

<sup>3</sup> *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.

8.3 *Combustion*:

8.3.1 Insert the stopper including the sample into the flask. Turn the flask upside down and fix it inside the combustion apparatus.

8.3.2 Lower the safety shield and start ignition by pushing the button.

8.3.3 After 5 min take the flask out of the apparatus and stir for 30 min with a magnetic stirrer.

8.3.4 Rinse the platinum grids with approximately 20 cm<sup>3</sup> of deionized water, while holding it above the combustion flask.

8.3.5 Transfer the liquid from the combustion flask to a 100 cm<sup>3</sup> volumetric flask, rinse, and make up to the mark with deionized water.

8.4 *Titration*:

8.4.1 Pipet 10 cm<sup>3</sup> from the volumetric flask into the titration beaker.

8.4.2 Add 40 cm<sup>3</sup> 2-propanol and 0.1 cm<sup>3</sup> Thorin solution.

NOTE 5—It is important to respect the proportion of water/2-propanol (1:4).

8.4.3 Degas in the ultrasonic bath for 5 min.

8.4.4 Add barium perchlorate until the color changes from yellow to pale rose. Preferably use an automatic titrator equipped with photoelectrical detector.

8.5 *Standardization of the Barium Perchlorate Solution*:

8.5.1 Pipet 10 cm<sup>3</sup> of sulfuric acid standard solution into a titration beaker and proceed as described in 8.4. Result = V<sub>0</sub>.

NOTE 6—Ion chromatography can be used as an alternative method for sulfate determination.

## 9. Calibration

9.1 Use sulfanilic acid as a standard for calibration. Carry out the test as described in Sections 7 and 8. The theoretical sulfur content of sulfanilic acid is 18.5 %. If this value is not obtained within ±0.2 % (absolute), check reagents and apparatus, glassware, and calculation.

## 10. Calculation

10.1 Calculate the sulfur content in percentage, according to the following equation:

$$S = \frac{V \cdot F \cdot 0.1603 \cdot V_{total}}{m \cdot V_{aliquot}} \cdot 100 \quad (1)$$

where:

V = volume of barium perchlorate solution,  
 F = factor of barium perchlorate solution  $F = 10.41/V_0$   
 (determined using sulfuric acid standard solution),

0.1603 = conversion factor from cm<sup>3</sup> 0.005 m barium perchlorate solution to mg S,

V<sub>total</sub> = total volume of sulfate-containing solution (here: 100 cm<sup>3</sup>),

V<sub>aliquot</sub> = aliquot part used for titration, cm<sup>3</sup>, and  
 m = original sample mass, mg.